

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction. Write the chemical formula(s) of the species involved in the reaction.

1. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
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18. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
19. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
20. Write the equation and/or the chemical formula(s) of the species involved in the reaction.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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Patrick Vollmar



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