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web solution 1 the two half reactions balanced as if in acidic solution $2\text{Cr}_2\text{O}_7 + 14\text{H}^+ + 6\text{Fe}^{2+} \rightarrow 4\text{Cr}^{3+} + 3\text{Fe}^{3+} + 7\text{H}_2\text{O}$ make electrons equal convert to basic solution $4\text{Cr}_2\text{O}_7 + 14\text{H}_2\text{O} + 6\text{Fe}^{2+} \rightarrow 4\text{Cr}^{3+} + 3\text{Fe}^{3+} + 14\text{H}_2\text{O}_2$ multiplied by a factor of 2 2e 2h2o 2o2 h2o2 2oh 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni²⁺ ni³⁺ a²⁺ a³⁺ q²⁻ e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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