

Chemistry Study Guide

Periodic Table of the Elements

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<div><div>Li</div><div>3</div><div>6.941</div></div>	<div><div>Be</div><div>4</div><div>9.012</div></div>																	<div><div>B</div><div>5</div><div>10.81</div></div>	<div><div>C</div><div>6</div><div>12.01</div></div>	<div><div>N</div><div>7</div><div>14.01</div></div>	<div><div>O</div><div>8</div><div>15.99</div></div>	<div><div>F</div><div>9</div><div>18.99</div></div>	<div><div>Ne</div><div>10</div><div>20.18</div></div>
<div><div>Na</div><div>11</div><div>22.99</div></div>	<div><div>Mg</div><div>12</div><div>24.31</div></div>																	<div><div>Al</div><div>13</div><div>26.98</div></div>	<div><div>Si</div><div>14</div><div>28.09</div></div>	<div><div>P</div><div>15</div><div>30.97</div></div>	<div><div>S</div><div>16</div><div>32.06</div></div>	<div><div>Cl</div><div>17</div><div>35.45</div></div>	<div><div>Ar</div><div>18</div><div>39.95</div></div>
<div><div>K</div><div>19</div><div>39.09</div></div>	<div><div>Ca</div><div>20</div><div>40.08</div></div>	<div><div>Sc</div><div>21</div><div>44.96</div></div>	<div><div>Ti</div><div>22</div><div>47.88</div></div>	<div><div>V</div><div>23</div><div>50.94</div></div>	<div><div>Cr</div><div>24</div><div>51.99</div></div>	<div><div>Mn</div><div>25</div><div>54.94</div></div>	<div><div>Fe</div><div>26</div><div>55.85</div></div>	<div><div>Co</div><div>27</div><div>58.93</div></div>	<div><div>Ni</div><div>28</div><div>58.71</div></div>	<div><div>Cu</div><div>29</div><div>63.55</div></div>	<div><div>Zn</div><div>30</div><div>65.38</div></div>	<div><div>Ga</div><div>31</div><div>69.72</div></div>	<div><div>Ge</div><div>32</div><div>72.64</div></div>	<div><div>As</div><div>33</div><div>74.92</div></div>	<div><div>Se</div><div>34</div><div>78.96</div></div>	<div><div>Br</div><div>35</div><div>79.90</div></div>	<div><div>Kr</div><div>36</div><div>83.80</div></div>						
<div><div>Rb</div><div>37</div><div>85.47</div></div>	<div><div>Sr</div><div>38</div><div>87.62</div></div>	<div><div>Y</div><div>39</div><div>88.91</div></div>	<div><div>Zr</div><div>40</div><div>91.22</div></div>	<div><div>Nb</div><div>41</div><div>92.91</div></div>	<div><div>Mo</div><div>42</div><div>95.94</div></div>	<div><div>Tc</div><div>43</div><div>98.00</div></div>	<div><div>Ru</div><div>44</div><div>101.07</div></div>	<div><div>Rh</div><div>45</div><div>102.91</div></div>	<div><div>Pd</div><div>46</div><div>106.42</div></div>	<div><div>Ag</div><div>47</div><div>107.87</div></div>	<div><div>Cd</div><div>48</div><div>112.41</div></div>	<div><div>In</div><div>49</div><div>114.82</div></div>	<div><div>Sn</div><div>50</div><div>118.71</div></div>	<div><div>Sb</div><div>51</div><div>121.76</div></div>	<div><div>Te</div><div>52</div><div>127.60</div></div>	<div><div>I</div><div>53</div><div>126.91</div></div>	<div><div>Xe</div><div>54</div><div>131.29</div></div>						
<div><div>Cs</div><div>55</div><div>132.91</div></div>	<div><div>Ba</div><div>56</div><div>137.33</div></div>																	<div><div>Tl</div><div>81</div><div>204.38</div></div>	<div><div>Pb</div><div>82</div><div>207.2</div></div>	<div><div>Bi</div><div>83</div><div>208.98</div></div>	<div><div>Po</div><div>84</div><div>209</div></div>	<div><div>At</div><div>85</div><div>210</div></div>	<div><div>Rn</div><div>86</div><div>222</div></div>
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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Standard Rate & Data Service



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