

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000?
 a. 1
 b. 1000
 c. 100
 d. 10

2. At standard $T_1 = 100^\circ\text{C}$, ΔH_{vap} of water is 40.7 kJ/mol. What is the enthalpy of vaporization?
 a. 40.7
 b. 40.7
 c. 40.7
 d. 40.7

3. The enthalpy of fusion is the lowest for:
 a. Water at 0°C
 b. Ice at 0°C
 c. Liquid water at 0°C
 d. Solid water at 0°C

4. For a process at $T = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, $\Delta S_{\text{vap}} = 109 \text{ J/mol}\cdot\text{K}$. What is the entropy change for the process?
 a. 109
 b. 40.7
 c. 109
 d. 40.7

5. What is the state of the water at 100°C and 1 atm?
 a. Liquid
 b. Solid
 c. Gas
 d. Superheated liquid

6. What is the boiling point of water?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

7. A process at 100°C and 1 atm, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, $\Delta S_{\text{vap}} = 109 \text{ J/mol}\cdot\text{K}$. What is the entropy change for the process?
 a. 109
 b. 40.7
 c. 109
 d. 40.7

8. For a process at $T = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, $\Delta S_{\text{vap}} = 109 \text{ J/mol}\cdot\text{K}$. What is the entropy change for the process?
 a. 109
 b. 40.7
 c. 109
 d. 40.7

9. At standard $T = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, $\Delta S_{\text{vap}} = 109 \text{ J/mol}\cdot\text{K}$. What is the entropy change for the process?
 a. 109
 b. 40.7
 c. 109
 d. 40.7

10. What is the boiling point of water?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

11. The enthalpy of fusion is the lowest for:
 a. Water at 0°C
 b. Ice at 0°C
 c. Liquid water at 0°C
 d. Solid water at 0°C

Multiple Choice Stoichiometry Test

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