

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|------------------------------------|
| acid | acid dissociation constant (K_a) | acid ionization constant (K_a) |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the K_a expression for the acid dissociation constant for each of the following acids. Write the K_b expression for the base dissociation constant for each of the following bases.

1. HNO_3 (strong acid)
2. HCl (strong acid)
3. H_2SO_4 (strong acid)
4. H_2SO_3 (weak acid)
5. H_2CO_3 (weak acid)
6. H_2O (weak acid)
7. H_2O (weak acid)
8. H_2O (weak acid)
9. H_2O (weak acid)
10. H_2O (weak acid)
11. H_2O (weak acid)
12. H_2O (weak acid)
13. H_2O (weak acid)
14. H_2O (weak acid)
15. H_2O (weak acid)
16. H_2O (weak acid)
17. H_2O (weak acid)
18. H_2O (weak acid)
19. H_2O (weak acid)
20. H_2O (weak acid)

1. OH^- (strong base)
2. OH^- (strong base)
3. OH^- (strong base)
4. OH^- (strong base)
5. OH^- (strong base)
6. OH^- (strong base)
7. OH^- (strong base)
8. OH^- (strong base)
9. OH^- (strong base)
10. OH^- (strong base)
11. OH^- (strong base)
12. OH^- (strong base)
13. OH^- (strong base)
14. OH^- (strong base)
15. OH^- (strong base)
16. OH^- (strong base)
17. OH^- (strong base)
18. OH^- (strong base)
19. OH^- (strong base)
20. OH^- (strong base)

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Linda Swisher, Gary A. Thibodeau



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