

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HI~~ **hydroiodic acid**
- ~~HNO₃~~ **nitric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sulfuric hydrosulfate~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~lithium hydroxide~~ **LiOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~strontium hydroxide~~ **Sr(OH)₂**

HClO₃ is NOT a strong acid & does not completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HI~~

~~HNO₃~~

~~Mg(OH)₂~~

HNO₂

NaOH

~~NaOH~~

H₂SO₄

3. What is the formula for hydronium? **H₃O⁺**

4. H₂C₂O₄ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: **$\text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{C}_2\text{O}_4^+ + \text{HC}_2\text{O}_4^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂C₂O₄ is a weak acid.**
- Would you define H₂C₂O₄ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂C₂O₄ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of H₂C₂O₄: **HC₂O₄⁻**
- Identify H₂C₂O₄ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H⁺ atoms.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F⁻: **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither: **neither.**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: **$\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO₄⁻: **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither: **amphoteric, because it has a negative charge (and can take a proton) and it has H⁺ atoms (and can lose one).**

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JL Elias



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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations

- 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
- 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
- 5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
 - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
- 6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
- 8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
- 9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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