

SOLIDS AND CRYSTAL STRUCTURE




Solids

- Molecules are held rigidly in position
- Molecules have no freedom of motion
- Less empty space in solids than liquids
- Virtually incompressible
- Molecules are arranged in regular configurations in three dimensions
- Has definite shape and volume

Intramolecular Forces

- hold atoms together in a molecule
- stabilize individual molecules
- involve chemical bonding


Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces

Force	Model	Basis of Attraction	Energy (kJ/mol)	Example
Bonding				
Ionic		Cation-anion	400–4000	NaCl
Covalent		Nuclei-shared e ⁻ pair	150–1100	H–H
Metallic		Cations-delocalized electrons	75–1000	Fe

Intermolecular Forces

- Attractive forces between molecules.
- Much weaker than intramolecular forces

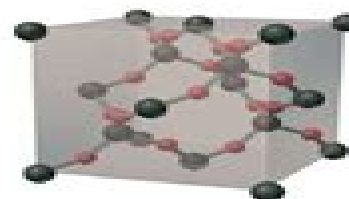
Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces

Nonbonding (Intermolecular)				
Ion-dipole		Ion-charge-dipole-charge	40–600	$\text{Na}^+ \cdots \text{O} \begin{array}{c} \text{H} \\ \diagup \diagdown \\ \text{H} \end{array}$
H bond		Polar bond to H-dipole-charge (high EN of N, O, F)	10–40	$\text{H}-\text{O}-\text{H} \cdots \text{O}-\text{H}$
Dipole-dipole		Dipole-charge	5–25	$\text{H}-\text{Cl} \cdots \text{Cl}-\text{H}$
Ion-induced dipole		Ion-charge-polarizable e ⁻ cloud	5–15	$\text{Na}^+ \cdots \text{O}_2$
Dipole-induced dipole		Dipole-charge-polarizable e ⁻ cloud	2–10	$\text{H}-\text{Cl} \cdots \text{Cl}-\text{Cl}$
Dispersion (London)		Polarizable e ⁻ clouds	0.05–40	$\text{F}-\text{F} \cdots \text{F}-\text{F}$

CATEGORIES OF SOLIDS

1. CRYSTALLINE SOLID

- It possesses rigid and long-range order.
- Its atoms, molecules or ions occupy specific positions and have regular structure.
- Ex: metals, many ceramics, some polymers.



2. AMORPHOUS SOLID

- Lacks regular three-dimensional arrangement of atom.
- without regular structure
- Ex: glass

The Crystal Structure Of Solids

Richard Bailey



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