

SOLID AND CRYSTAL STRUCTURE

Solids

- Molecules are held rigidly in position
- Molecules have no freedom of motion
- Less empty space in solids than liquids
- Virtually incompressible
- Molecules are arranged in regular configurations in three dimensions
- Has definite shape and volume

Intramolecular Forces

- hold atoms together in a molecule
- stabilize individual molecules
- involve chemical bonding

Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces

Force	Model	Basis of Attraction	Energy (kJ/mol)	Example
Bonding				
Ionic		Cation-anion	400-4000	NaCl
Covalent		Nuclei-shared e- pair	150-1100	H-H
Metallic		Cations-delocalized electrons	75-1000	Fe

Intermolecular Forces

- Attractive forces between molecules
- Much weaker than intramolecular forces

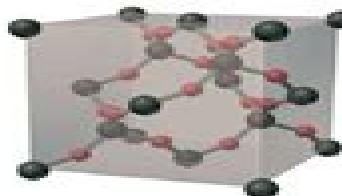
Table 12.2 Comparison of Bonding and Nonbonding (Intermolecular) Forces

Nonbonding (Intermolecular)				
Ion-dipole		Ion charge-dipole charge	40-400	$\text{Na}^+ \cdots \text{O}_2^-$
H bond		Polar bond to H-dipole charge (high EN of N, O, F)	30-40	$\text{H}-\text{O} \cdots \text{O}_2^-$
Disp-dipole		Dipole charges	3-25	$\text{H}-\text{O} \cdots \text{O}_2^-$
Ion-induced dipole		Ion charge-polarizable e- cloud	3-15	$\text{Na}^+ \cdots \text{O}_2^-$
Disp-induced dipole		Dipole charge-polarizable e- cloud	2-10	$\text{H}-\text{O} \cdots \text{O}_2^-$
Dispersion (London)		Polarizable e- clouds	0.05-40	$\text{F}-\text{F} \cdots \text{F}-\text{F}$

CATEGORIES OF SOLIDS

1. CRYSTALLINE SOLID

- It possesses rigid and long-range order.
- Its atoms, molecules or ions occupy specific positions and have regular structure.
- Ex: metals, many ceramics, some polymers.



2. AMORPHOUS SOLID

- Lacks regular three-dimensional arrangement of atom.
- without regular structure
- Ex: glass

The Crystal Structure Of Solids

Richard Bailey

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