

Chemistry Study Guide

Periodic Table of the Elements

<small>1</small> H <small>1.008</small>																	<small>2</small> He <small>4.003</small>																															
<small>3</small> Li <small>6.941</small>	<small>4</small> Be <small>9.012</small>																	<small>5</small> B <small>10.811</small>	<small>6</small> C <small>12.011</small>	<small>7</small> N <small>14.007</small>	<small>8</small> O <small>15.999</small>	<small>9</small> F <small>18.998</small>	<small>10</small> Ne <small>20.180</small>																									
<small>11</small> Na <small>22.990</small>	<small>12</small> Mg <small>24.305</small>																	<small>13</small> Al <small>26.982</small>	<small>14</small> Si <small>28.086</small>	<small>15</small> P <small>30.974</small>	<small>16</small> S <small>32.065</small>	<small>17</small> Cl <small>35.453</small>	<small>18</small> Ar <small>39.948</small>																									
<small>19</small> K <small>39.098</small>	<small>20</small> Ca <small>40.078</small>	<small>21</small> Sc <small>44.956</small>	<small>22</small> Ti <small>47.88</small>	<small>23</small> V <small>50.942</small>	<small>24</small> Cr <small>52.004</small>	<small>25</small> Mn <small>54.938</small>	<small>26</small> Fe <small>55.845</small>	<small>27</small> Co <small>58.933</small>	<small>28</small> Ni <small>58.693</small>	<small>29</small> Cu <small>63.546</small>	<small>30</small> Zn <small>65.38</small>	<small>31</small> Ga <small>69.723</small>	<small>32</small> Ge <small>72.64</small>	<small>33</small> As <small>74.922</small>	<small>34</small> Se <small>78.96</small>	<small>35</small> Br <small>79.904</small>	<small>36</small> Kr <small>83.80</small>																															
<small>37</small> Rb <small>85.468</small>	<small>38</small> Sr <small>87.62</small>	<small>39</small> Y <small>88.906</small>	<small>40</small> Zr <small>91.224</small>	<small>41</small> Nb <small>92.906</small>	<small>42</small> Mo <small>95.94</small>	<small>43</small> Tc <small>98.906</small>	<small>44</small> Ru <small>101.07</small>	<small>45</small> Rh <small>102.91</small>	<small>46</small> Pd <small>106.42</small>	<small>47</small> Ag <small>107.87</small>	<small>48</small> Cd <small>112.41</small>	<small>49</small> In <small>114.82</small>	<small>50</small> Sn <small>118.71</small>	<small>51</small> Sb <small>121.76</small>	<small>52</small> Te <small>127.6</small>	<small>53</small> I <small>126.91</small>	<small>54</small> Xe <small>131.29</small>																															
<small>55</small> Cs <small>132.91</small>	<small>56</small> Ba <small>137.33</small>			<small>57</small> Hf <small>178.49</small>	<small>58</small> Ta <small>180.95</small>	<small>59</small> W <small>183.84</small>	<small>60</small> Re <small>186.21</small>	<small>61</small> Os <small>190.23</small>	<small>62</small> Ir <small>192.22</small>	<small>63</small> Pt <small>195.08</small>	<small>64</small> Au <small>196.97</small>	<small>65</small> Hg <small>200.59</small>	<small>66</small> Tl <small>204.38</small>	<small>67</small> Pb <small>207.2</small>	<small>68</small> Bi <small>208.98</small>	<small>69</small> Po <small>209</small>	<small>70</small> At <small>210</small>	<small>71</small> Rn <small>222</small>																														
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Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr																																		

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎📌

Matter:

Anything that has **mass** and **occupies space** ✎🌐

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

Zhe Xu



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