

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the problem. Express answers in scientific notation when appropriate.

1. What substance has $[\text{H}^+] = 1.0 \times 10^{-7}$ moles/liter hydrogen ion?
2. What substance has $\text{pH} = 7$? The concentration of this acidic substance is 1.0×10^{-7} moles/liter.
3. Hydrochloric acid has a pH of 1.0 or 1.0×10^{-1} moles/liter. What is the concentration of this acid in moles/liter?
4. What is the pH of a solution that is 1.0×10^{-2} moles/liter?
5. What is the pH of a solution that is 1.0×10^{-12} moles/liter?
6. What is the pH of a solution that is 1.0×10^{-14} moles/liter?
7. What is the pH of a solution that is 1.0×10^{-14} moles/liter?

8. 1.0×10^{-7}
9. 1.0×10^{-7}
10. 1.0×10^{-1}
11. 1.0×10^{-1}
12. 1.0×10^{-1}
13. 1.0×10^{-1}
14. 1.0×10^{-1}
15. 1.0×10^{-1}
16. 1.0×10^{-1}
17. 1.0×10^{-1}
18. 1.0×10^{-1}

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S Ashworth



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