

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a and K_b | 6. relative acidity |
| 7. ionization constant | 8. weak solution | 9. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include proper units when appropriate.

1. Water autoionizes to H^+ and hydroxide ions. The concentration of H^+ in the autoionization of water at 25°C is 1.0×10^{-7} mol/L.
2. The autoionization constant for water is 1.0×10^{-14} at 25°C.
3. The autoionization constant for water is 1.0×10^{-14} at 25°C.
4. The autoionization constant for water is 1.0×10^{-14} at 25°C.
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20. The autoionization constant for water is 1.0×10^{-14} at 25°C.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. 1.0×10^{-14}
3. 1.0×10^{-14}
4. 1.0×10^{-14}
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19 2 Hydrogen Ions And Acidity Answer Key

Valerie C. Scanlon, Tina Sanders



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