

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. weak acid | 6. conjugate base |
| | 7. weak base | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constants involved that are indicated in the spaces provided. Be sure to include proper units when appropriate.

1. Water autoionizes to H^+ and hydroxide ions. The concentration of H^+ in the autoionization of water at 25°C is $1.0 \times 10^{-7} \text{ M}$.
2. The K_a of acetic acid is 1.8×10^{-5} .
3. The K_b of ammonia is 1.8×10^{-5} .
4. The K_a of hydrofluoric acid is 6.6×10^{-4} .
5. The K_b of the conjugate base of acetic acid is 5.6×10^{-10} .
6. The K_a of the conjugate base of ammonia is 5.6×10^{-10} .
7. The K_a of the conjugate base of hydrofluoric acid is 1.5×10^{-3} .

8. $\text{pH} = 7.0$
9. $\text{pH} = 7.0$
10. $\text{pH} = 7.0$
11. $\text{pH} = 7.0$
12. $\text{pH} = 7.0$
13. $\text{pH} = 7.0$
14. $\text{pH} = 7.0$
15. $\text{pH} = 7.0$
16. $\text{pH} = 7.0$
17. $\text{pH} = 7.0$
18. $\text{pH} = 7.0$

19 2 Hydrogen Ions And Acidity Answer Key

Anthony S. Fauci



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