

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                |          |          |
|----------------|----------|----------|
| acid           | 1. $K_a$ | 2. $K_b$ |
| base           | 3. $K_a$ | 4. $K_b$ |
| conjugate acid | 5. $K_a$ | 6. $K_b$ |
| conjugate base | 7. $K_a$ | 8. $K_b$ |

## Key Equations

- |                                    |                                     |
|------------------------------------|-------------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| 3. $K_a = 10^{-pH}$                | 4. $K_b = 10^{-pOH}$                |

## Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing double bonds where necessary. Write the chemical formula for each of the following acids, bases, and salts.

1.  $K_a$  of  $CH_3COOH$  is  $1.8 \times 10^{-5}$ .
2.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ .
3.  $K_a$  of  $CH_3COOH$  is  $1.8 \times 10^{-5}$ .
4.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ .
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8.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ .
9.  $K_a$  of  $CH_3COOH$  is  $1.8 \times 10^{-5}$ .
10.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ .

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