



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log \frac{1}{[\text{H}^+]}$

$$\text{pH} = -\log[\text{H}^+]$$

Practice Problems

1. Calculate the pH of a 0.001 M HCl solution.

Answers

1. 3.00	2. 1.00	3. 1.00
4. 1.00	5. 1.00	6. 1.00
7. 1.00	8. 1.00	9. 1.00
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55. 1.00	56. 1.00	57. 1.00
58. 1.00	59. 1.00	60. 1.00
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