



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acid	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acid
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Problems

Use the following equations to calculate the pH of the following acids. The concentration of each acid is $1.0 \times 10^{-3} \text{ M}$.

1. H_2SO_4 (dilute sulfuric acid) $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
concentration: $1.0 \times 10^{-3} \text{ M}$ concentration of H^+ is $2.0 \times 10^{-3} \text{ M}$

2. H_3PO_4 (dilute phosphoric acid) $\text{H}_3\text{PO}_4 \rightarrow 3\text{H}^+ + \text{PO}_4^{3-}$
concentration: $1.0 \times 10^{-3} \text{ M}$ concentration of H^+ is $3.0 \times 10^{-3} \text{ M}$

3. HCl (dilute hydrochloric acid) $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$
concentration: $1.0 \times 10^{-3} \text{ M}$ concentration of H^+ is $1.0 \times 10^{-3} \text{ M}$

Acid	Concentration	pH
H_2SO_4	$1.0 \times 10^{-3} \text{ M}$	2.0
H_3PO_4	$1.0 \times 10^{-3} \text{ M}$	2.0
HCl	$1.0 \times 10^{-3} \text{ M}$	2.0
H_2O	$1.0 \times 10^{-7} \text{ M}$	7.0
NaOH	$1.0 \times 10^{-13} \text{ M}$	13.0
H_2O	$1.0 \times 10^{-14} \text{ M}$	14.0

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JE Gale

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