

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$ | $pOH = -\log[OH^-]$ |

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, and give the name of each compound.

1. $K_a = 1.0 \times 10^{-4}$ (weak acid)

2. $K_b = 1.0 \times 10^{-4}$ (weak base)

3. $K_a = 1.0 \times 10^{-4}$ (weak acid)

4. $K_b = 1.0 \times 10^{-4}$ (weak base)

5. $K_a = 1.0 \times 10^{-4}$ (weak acid)

6. $K_b = 1.0 \times 10^{-4}$ (weak base)

7. $K_a = 1.0 \times 10^{-4}$ (weak acid)

8. $K_b = 1.0 \times 10^{-4}$ (weak base)

9. $K_a = 1.0 \times 10^{-4}$ (weak acid)

10. $K_b = 1.0 \times 10^{-4}$ (weak base)

11. $K_a = 1.0 \times 10^{-4}$ (weak acid)

1. $K_a = 1.0 \times 10^{-4}$
2. $K_b = 1.0 \times 10^{-4}$
3. $K_a = 1.0 \times 10^{-4}$
4. $K_b = 1.0 \times 10^{-4}$
5. $K_a = 1.0 \times 10^{-4}$
6. $K_b = 1.0 \times 10^{-4}$
7. $K_a = 1.0 \times 10^{-4}$
8. $K_b = 1.0 \times 10^{-4}$
9. $K_a = 1.0 \times 10^{-4}$
10. $K_b = 1.0 \times 10^{-4}$
11. $K_a = 1.0 \times 10^{-4}$

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