



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>hydrogen ion</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li><li>pH</li></ul>	<ul style="list-style-type: none"><li>hydrogen ion concentration</li><li>acidic</li><li>basic</li><li>neutral</li><li>titration</li><li>titrant</li></ul>
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### Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log[\text{OH}^-]$

$$\text{pH} = -\log[\text{H}^+]$$

### Principles of Acids and Bases

Acids and bases are substances that have the ability to ionize in water to produce hydrogen ions and hydroxide ions, respectively.

Hydrogen ions ( $\text{H}^+$ ) are the smallest hydrogen atoms. They are positively charged because they have lost an electron. They are also called protonic acids because they contain protons.

Hydroxide ions ( $\text{OH}^-$ ) are negatively charged because they have an extra electron. They are also called hydroxide bases because they contain hydroxide ions.

Acids are substances that have the ability to ionize in water to produce hydrogen ions ( $\text{H}^+$ ). Acids are also called protonic acids because they contain protons.

Bases are substances that have the ability to ionize in water to produce hydroxide ions ( $\text{OH}^-$ ). Bases are also called hydroxide bases because they contain hydroxide ions.

### Acid-Base Reactions

1. Neutralization	2. Acid-Base Titration
3. Acid-Base Indicators	4. Buffers
5. Dissociation	6. Ionization
7. Ion Exchange	8. Ion-Pair Formation
9. Salts	10. Polyacids and Polybases
11. Polyprotic Acids	12. Polyprotic Bases
13. Polyacid-Polybase	14. Polyacid-Polyacid
15. Polybase-Polybase	16. Polyacid-Polybase
17. Polyacid-Polyacid	18. Polybase-Polyacid
19. Polyacid-Polyacid	20. Polybase-Polybase
21. Polyacid-Polyacid	22. Polybase-Polyacid
23. Polyacid-Polyacid	24. Polybase-Polyacid
25. Polyacid-Polyacid	26. Polybase-Polyacid
27. Polyacid-Polyacid	28. Polybase-Polyacid
29. Polyacid-Polyacid	30. Polybase-Polyacid
31. Polyacid-Polyacid	32. Polybase-Polyacid
33. Polyacid-Polyacid	34. Polybase-Polyacid
35. Polyacid-Polyacid	36. Polybase-Polyacid
37. Polyacid-Polyacid	38. Polybase-Polyacid
39. Polyacid-Polyacid	40. Polybase-Polyacid
41. Polyacid-Polyacid	42. Polybase-Polyacid
43. Polyacid-Polyacid	44. Polybase-Polyacid
45. Polyacid-Polyacid	46. Polybase-Polyacid
47. Polyacid-Polyacid	48. Polybase-Polyacid
49. Polyacid-Polyacid	50. Polybase-Polyacid
51. Polyacid-Polyacid	52. Polybase-Polyacid
53. Polyacid-Polyacid	54. Polybase-Polyacid
55. Polyacid-Polyacid	56. Polybase-Polyacid
57. Polyacid-Polyacid	58. Polybase-Polyacid
59. Polyacid-Polyacid	60. Polybase-Polyacid
61. Polyacid-Polyacid	62. Polybase-Polyacid
63. Polyacid-Polyacid	64. Polybase-Polyacid
65. Polyacid-Polyacid	66. Polybase-Polyacid
67. Polyacid-Polyacid	68. Polybase-Polyacid
69. Polyacid-Polyacid	70. Polybase-Polyacid
71. Polyacid-Polyacid	72. Polybase-Polyacid
73. Polyacid-Polyacid	74. Polybase-Polyacid
75. Polyacid-Polyacid	76. Polybase-Polyacid
77. Polyacid-Polyacid	78. Polybase-Polyacid
79. Polyacid-Polyacid	80. Polybase-Polyacid
81. Polyacid-Polyacid	82. Polybase-Polyacid
83. Polyacid-Polyacid	84. Polybase-Polyacid
85. Polyacid-Polyacid	86. Polybase-Polyacid
87. Polyacid-Polyacid	88. Polybase-Polyacid
89. Polyacid-Polyacid	90. Polybase-Polyacid
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95. Polyacid-Polyacid	96. Polybase-Polyacid
97. Polyacid-Polyacid	98. Polybase-Polyacid
99. Polyacid-Polyacid	100. Polybase-Polyacid

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## **Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings
  - 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks

- 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
- 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
- 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options

6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats

- ePUB, PDF, MOBI, and More
- 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
- 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
- Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
- Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key

8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key

9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key

- Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key

- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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