

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

## Vocabulary

- |                   |   |                     |
|-------------------|---|---------------------|
| 1. autoionization | 2. hydronium ion ( $\text{H}_3\text{O}^+$ ) | 3. acid solution    |
| 4. autoionization | 5. acid rain                                | 6. neutral solution |
|                   | 7. acid solution                            | 8. pH               |

## Key Equation

- |                                     |   |
|-------------------------------------|---|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the missing quantities that are indicated in the partial equations below. Complete each item completely.

1. What substance is \_\_\_\_\_, which hydrogen ion concentration is \_\_\_\_\_ the concentration of that which produces an acidic solution? \_\_\_\_\_
2. Hydrochloric acid has a pH of \_\_\_\_\_, so it is \_\_\_\_\_ acidic. \_\_\_\_\_ concentration of that acid is \_\_\_\_\_.
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# 19 2 Hydrogen Ions And Acidity Answer Key

**Teresa Audesirk, Gerald  
Audesirk, Bruce E. Byers**



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