

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the ionization constant of water. Write the equation and K_w at the temperature of 25°C.
2. Hydrochloric acid dissociates from HCl to H^+ and Cl^- . Write the K_a expression and K_a value for this acid.
3. Acetic acid, CH_3COOH , dissociates to CH_3COO^- and H^+ . Write the K_a expression and K_a value for this acid.
4. OH^- is the conjugate base of water. Write the K_b expression and K_b value for this base.
5. CH_3COO^- is the conjugate base of acetic acid. Write the K_b expression and K_b value for this base.

6. K_a for HCl is 1.0×10^7 .
7. K_b for OH^- is 1.0×10^{-14} .
8. K_a for HCl is 1.0×10^7 .
9. K_b for OH^- is 1.0×10^{-14} .
10. K_a for HCl is 1.0×10^7 .
11. K_b for OH^- is 1.0×10^{-14} .
12. K_a for HCl is 1.0×10^7 .
13. K_b for OH^- is 1.0×10^{-14} .
14. K_a for HCl is 1.0×10^7 .
15. K_b for OH^- is 1.0×10^{-14} .
16. K_a for HCl is 1.0×10^7 .
17. K_b for OH^- is 1.0×10^{-14} .
18. K_a for HCl is 1.0×10^7 .

19 2 Hydrogen Ions And Acidity Answer Key

RS Peters



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