

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak bases and acids.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak bases and acids.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- acid dissociation constant
- base dissociation constant

- acid dissociation constant
- base dissociation constant
- acid dissociation constant

- acid dissociation constant
- base dissociation constant
- acid dissociation constant

Key Equations

- 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)
- 2. $K_b = \frac{[OH^-][BH^+]}{[B]}$ (Equation 19.2)

- 3. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)
- 4. $K_b = \frac{[OH^-][BH^+]}{[B]}$ (Equation 19.2)

Part A Completion

Write the chemical formula for the acid or base corresponding to the K_a or K_b value given. Write the chemical formula for the acid or base corresponding to the K_a or K_b value given.

1. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

2. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

3. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

4. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

5. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

6. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

7. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

8. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

9. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

10. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

11. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

12. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

13. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

14. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

15. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

16. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

17. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

18. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

19. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

20. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

21. $K_a = 1.8 \times 10^{-4}$ (Equation 19.1)

22. $K_b = 1.8 \times 10^{-4}$ (Equation 19.2)

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