

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify ionization levels based on the pH of acids and bases.
2. Calculate hydrogen ion concentration, molar concentration, and pH.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-------------------------------|-----------------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid ionization constant |
| 4. autoionization | 5. molar concentration | 6. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration, molar concentration, or the molar concentration of the acid or base in the space provided. Round answers to the correct number of significant figures.

1. What is the $[\text{H}^+]$ of a solution with a pH of 4.5?
2. What is the $[\text{H}^+]$ of a solution with a pH of 10.2?
3. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
4. What is the $[\text{H}^+]$ of a solution with a pH of 12.0?
5. What is the $[\text{H}^+]$ of a solution with a pH of 1.5?
6. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
7. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
8. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
9. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
10. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
11. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
12. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
13. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
14. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
15. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
16. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
17. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
18. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
19. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?
20. What is the $[\text{H}^+]$ of a solution with a pH of 1.0?

1. 1.0×10^{-4}
2. 1.0×10^{-10}
3. 1.0×10^{-1}
4. 1.0×10^{-12}
5. 1.0×10^{-1}
6. 1.0×10^{-1}
7. 1.0×10^{-1}
8. 1.0×10^{-1}
9. 1.0×10^{-1}
10. 1.0×10^{-1}
11. 1.0×10^{-1}
12. 1.0×10^{-1}
13. 1.0×10^{-1}
14. 1.0×10^{-1}
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16. 1.0×10^{-1}
17. 1.0×10^{-1}
18. 1.0×10^{-1}
19. 1.0×10^{-1}
20. 1.0×10^{-1}

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M Woodhall



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