

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- acid
- base

- acidic solution
- basic solution

- acidic solution
- basic solution

Key Equations

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad (19.1)$$

$$K_b = \frac{[OH^-][BH^+]}{[B]} \quad (19.2)$$

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, or salt.

1. $K_a = 1.0 \times 10^{-4}$ (weak acid)

2. $K_b = 1.0 \times 10^{-4}$ (weak base)

3. $K_a = 1.0 \times 10^{-4}$ (weak acid)

4. $K_b = 1.0 \times 10^{-4}$ (weak base)

5. $K_a = 1.0 \times 10^{-4}$ (weak acid)

6. $K_b = 1.0 \times 10^{-4}$ (weak base)

7. $K_a = 1.0 \times 10^{-4}$ (weak acid)

8. $K_b = 1.0 \times 10^{-4}$ (weak base)

9. $K_a = 1.0 \times 10^{-4}$ (weak acid)

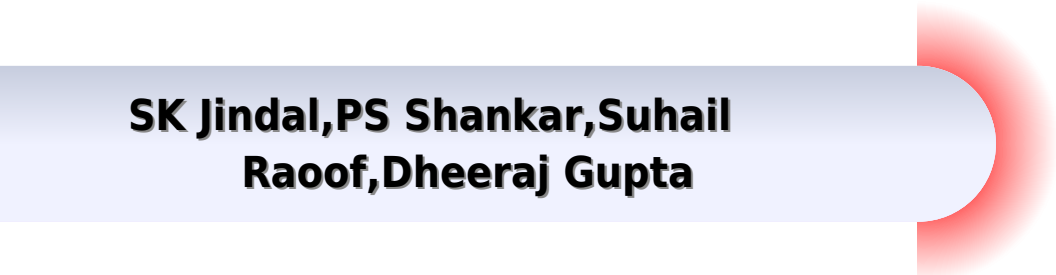
10. $K_b = 1.0 \times 10^{-4}$ (weak base)

11. $K_a = 1.0 \times 10^{-4}$ (weak acid)

1. $K_a = 1.0 \times 10^{-4}$
2. $K_b = 1.0 \times 10^{-4}$
3. $K_a = 1.0 \times 10^{-4}$
4. $K_b = 1.0 \times 10^{-4}$
5. $K_a = 1.0 \times 10^{-4}$
6. $K_b = 1.0 \times 10^{-4}$
7. $K_a = 1.0 \times 10^{-4}$
8. $K_b = 1.0 \times 10^{-4}$
9. $K_a = 1.0 \times 10^{-4}$
10. $K_b = 1.0 \times 10^{-4}$
11. $K_a = 1.0 \times 10^{-4}$

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