

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the equilibrium constant for autoionization. K_w is the concentration of those which participate in the reaction. K_w is 1.0×10^{-14} .
2. The equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ is K_w . K_w is the concentration of those which participate in the reaction. K_w is 1.0×10^{-14} .
3. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$. K_w is the concentration of those which participate in the reaction. K_w is 1.0×10^{-14} .

4. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
5. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
6. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
7. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
8. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
9. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
10. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
11. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
12. K_w is the equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.

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Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key

8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key
9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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