

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, from pK_a and pK_b values.
2. Calculate pH and pOH for weak acids and bases, and vice versa, from K_a and K_b values.
3. Calculate the percent ionization of weak acids.

Vocabulary

acid dissociation

acid ionization

acid ionization constant (K_a)

acid strength

acid–base indicator

base dissociation

base ionization

pH

Key Equations

$K_a = 10^{-pK_a}$ (Equation 19.2) or $pK_a = -\log K_a$ (Equation 19.3)

$K_b = 10^{-pK_b}$ (Equation 19.4) or $pK_b = -\log K_b$ (Equation 19.5)

Part A Completion

Write the appropriate formula or name for each compound, state whether it is a strong or weak acid or base, and write the K_a or K_b value for each acid or base.

1. Strong acid: HNO_3 (nitric acid) $K_a = 10^{-1.4}$

2. Weak acid: H_2CO_3 (carbonic acid) $K_a = 4.3 \times 10^{-7}$

3. Strong base: $NaOH$ (sodium hydroxide) $K_b = 10^{-14}$

4. Weak base: NH_3 (ammonia) $K_b = 1.8 \times 10^{-5}$

5. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

6. Weak acid: H_2SO_4 (sulfuric acid) $K_a = 1.0 \times 10^{-2}$

7. Strong base: KOH (potassium hydroxide) $K_b = 10^{-14}$

8. Weak acid: $H_2PO_4^-$ (dihydrogen phosphate) $K_a = 6.3 \times 10^{-8}$

9. Strong acid: HNO_3 (nitric acid) $K_a = 10^{-1.4}$

10. Weak base: NH_3 (ammonia) $K_b = 1.8 \times 10^{-5}$

11. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

12. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

13. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

14. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

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18. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

19. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

20. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

21. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

22. Strong acid: HCl (hydrochloric acid) $K_a = 10^{-1.1}$

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