

## 14.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
| 7. autoionization | 8. acid rain                | 9. acid          |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water.
3. Write the equation for the autoionization of water.
4. Write the equation for the autoionization of water.
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18. Write the equation for the autoionization of water.
19. Write the equation for the autoionization of water.
20. Write the equation for the autoionization of water.

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
5.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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19.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
20.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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**Mark Steyvers**



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## **Table of Contents 19 2 Hydrogen Ions And Acidity Work Sheet Answers**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers User Reviews and Ratings
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers and Bestseller Lists
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  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers eBook Subscription Services
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Budget-Friendly Options

6. Navigating 19 2 Hydrogen Ions And Acidity Work Sheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Interactive Elements 19 2 Hydrogen Ions And Acidity Work Sheet Answers
8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Work Sheet Answers
9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Work Sheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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