

Physical Chemistry Formula Revision

Basic Concepts

$$\text{Average atomic mass} = \frac{a_1 \cdot n_1 + a_2 \cdot n_2 + a_3 \cdot n_3}{100}$$

a_1, a_2, a_3 ... = atomic mass of Isotopes

n_1, n_2, n_3 ... = percentage abundance

Mole concept

$$\text{no. of moles, } n = \frac{\text{given mass}}{\text{molar mass}} = \frac{m}{M}$$

$$n = \frac{\text{no. of particles}}{\text{Avogadro's no.}} = \frac{N}{6.023 \times 10^{23}}$$

$$n = \frac{\text{vol. of gas}}{\text{vol. of gas at STP}} = \frac{V}{22.41}$$

Concentration Terms

$$\text{Percentage, } \% = \frac{W_b \text{ or } V_b}{W \text{ or } V} \times 100$$

$$\text{Molarity, } M = \frac{n_b}{V(l)} = \frac{W_b \times 1000}{M_b \times V(\text{ml})}$$

$$\text{Normality, } N = \frac{n_b}{V(l)} = \frac{W_b \times 1000}{M_b \times V(\text{ml})} \times \frac{L_p}{M_p} \times \frac{1000}{n_p}$$

$$\text{parts per million, ppm} = \frac{W_b}{W} \times 10^6$$

$$\text{Molality, } m = \frac{n_b}{W_b(\text{kg})} = \frac{W_b \times 1000}{M_b \times W_b(\text{kg})}$$

$$\text{Mole fraction, } x_b = \frac{n_b}{n_a + n_b + n_c}$$

Stoichiometry



$$\frac{\text{moles of A}}{a} = \frac{\text{moles of B}}{b} = \frac{\text{moles of C}}{c}$$

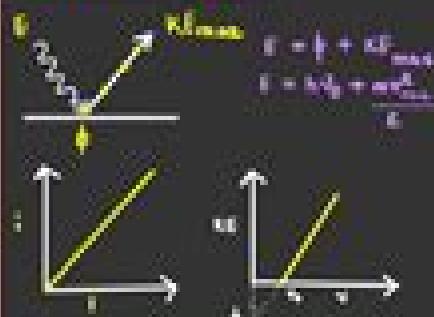
Limiting Ag: $\frac{\text{given amt.}}{\text{lowest stoichiometric coeff.}}$ \rightarrow lowest

Atomic Structure

Planck's Quantum Theory

$$E = hc \frac{1}{\lambda} \quad E(\text{eV}) = \frac{1240}{\lambda(\text{nm})}$$

Photo-Electric Effect



De-Broglie's Hypothesis

$$\omega^2 = \frac{h^2}{\lambda^2} \rightarrow \lambda = \frac{h}{\sqrt{2mE}}$$

Schrodinger's model

$$\psi = \frac{1}{\sqrt{2\pi}} \times \frac{e^{i\theta}}{r} \times \frac{e^{i\phi}}{\sqrt{2}} \times \frac{e^{i\theta}}{2} \times \frac{e^{i\phi}}{2}$$

$$E = -\frac{1}{2} \times \frac{e^2}{r^2} \times V \quad T_E = \frac{p^2}{2m} \times -0.6$$

$$\frac{1}{r} = k_1 \cdot r^2 \left[\frac{1}{r_1^2} + \frac{1}{r_2^2} \right] \quad k_1 = 1.13 \times 10^{-10} \text{ m}^{-1}$$

no. no. of orbital lines when 2ⁿ atoms from $n_1, n_2, n_3, \dots, n_n$ \rightarrow $(n_1 + 1)(n_2 + 1) \dots (n_n + 1)$

Heisenberg Uncertainty Principle: $\Delta x \Delta p \geq \frac{h}{4\pi}$

Quantum Model

total nodes = n-1

angular nodes = l

radial nodes = n-1-l

Angular momentum

$$z\text{-orbital} = \sqrt{1 + 3\cos^2 \theta} / 2$$

$$p\text{-orbital} = \sqrt{1 + 3\cos^2 \theta} / 2$$

$$\text{spin} = \sqrt{1 + 3\cos^2 \theta} / 2$$

The Nature Of Physical Chemistry

**Lawrence Stamper Darken, Robert
Wilton Gurry**

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The Nature of Physical Chemistry Dudley J. Luke,1969 [The Nature of Physical Chemistry](#) ,19?? [Principles of Physical Chemistry](#) Hans Kuhn,Horst-Dieter Försterling,David H. Waldeck,2009-03-17 Principles of Physical Chemistry Second Edition uniquely uses simple physical models as well as rigorous treatments for understanding molecular and supramolecular systems and processes In this way the presentation assists students in developing an intuitive understanding of the subjects as well as skill in quantitative manipulations The unifying nature of physical chemistry is emphasized in the book by its organization beginning with atoms and molecules and proceeding to molecular assemblies of increasing complexity ending with the emergence of matter that carries information i e the origin of life a physicochemical process of unique importance The aim is to show the broad scope and coherence of physical chemistry [The Nature of Physical Chemistry](#), by Robert Ruffnell and Dudley J. Luke Robert Tuffnell, *An Introduction to the Principles of Physical Chemistry from the Standpoint of Modern Atomistics and Thermo-dynamics* Edward Wight Washburn,1915 **Physical Chemistry** Kenneth S Schmitz,2016-11-11 Physical Chemistry Concepts and Theory provides a comprehensive overview of physical and theoretical chemistry while focusing on the basic principles that unite the sub disciplines of the field With an emphasis on multidisciplinary as well as interdisciplinary applications the book extensively reviews fundamental principles and presents recent research to help the reader make logical connections between the theory and application of physical chemistry concepts Also available from the author Physical Chemistry Multidisciplinary Applications ISBN 9780128005132 Describes how materials behave and chemical reactions occur at the molecular and atomic levels Uses theoretical constructs and mathematical computations to explain chemical properties and describe behavior of molecular and condensed matter Demonstrates the connection between math and chemistry and how to use math as a powerful tool to predict the properties of chemicals Emphasizes the intersection of chemistry math and physics and the resulting applications across many disciplines of science [Essential Results in Chemical Physics and Physical Chemistry](#) Alexander N. Goloschapov,Gennadii Efremovich Zaikov,Vladimir Viktorovich Ivanov,2005 The contents included in this book are Preface Spin Probes for the Study of Intact and Cancer Cell Membranes Sulphur as a Stabiliser of Polyvinylchloride Universality of Free Energies Linearity Principle in Solution Chemistry The KBr Action on the rate of H₂O₂ Decomposition in Alkaline Medium Fireproof Materials containing Nanostructures Principles of Formation Fireproof Intumescent Coating Foamcoke Structure Regulation by Carbon Metal containing Nanostructures Upholstery Fire Barriers based on Natural Fibres Structural Criterion on Change of a Kinetic Curves Type in the Process of a Thermooxidative Degradation and Alternative View at the Universe It also includes Effect of the Cationic Polyelectrolyte Molecular Mass on the Flocculation Kinetics and the Efficiency of Polymer Precipitation from Latexes Co polymers with Cyclic Fragments in Dimethylsiloxane Backbone O Fractal Physics of the Polycondensation Processes The Problem of Structural Physical Organisation of Polymeric Non Crystalline Phase and

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