

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element within each molecule or formula.

1. Acetic acid, CH_3COOH
2. Hydrofluoric acid, HF
3. Hydrochloric acid, HCl
4. Hydrobromic acid, HBr
5. Hydroiodic acid, HI
6. Nitric acid, HNO_3
7. Sulfuric acid, H_2SO_4
8. Phosphoric acid, H_3PO_4
9. Acetic acid, CH_3COOH
10. Hydrofluoric acid, HF
11. Hydrochloric acid, HCl
12. Hydrobromic acid, HBr
13. Hydroiodic acid, HI
14. Nitric acid, HNO_3
15. Sulfuric acid, H_2SO_4
16. Phosphoric acid, H_3PO_4

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14. Nitric acid, HNO_3
15. Sulfuric acid, H_2SO_4
16. Phosphoric acid, H_3PO_4

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