

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) indicated. Show the calculation for pH or percent dissociation when required. Assume acids are monoprotic unless noted.

1. What substance is H^+ in water? Hydrogen ion is H_3O^+ .
2. Write the equation for the autoionization of water which produces an H^+ ion and a hydroxide ion. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. What is the pH of a 0.001 M solution of HCl in water? $\text{pH} = 3$
4. What is the pH of a 0.001 M solution of NaOH in water? $\text{pH} = 11$
5. What is the pH of a 0.001 M solution of H_2SO_4 in water? $\text{pH} = 1$
6. What is the pH of a 0.001 M solution of NaOH in water? $\text{pH} = 11$
7. What is the pH of a 0.001 M solution of H_2SO_4 in water? $\text{pH} = 1$
8. What is the pH of a 0.001 M solution of NaOH in water? $\text{pH} = 11$
9. What is the pH of a 0.001 M solution of H_2SO_4 in water? $\text{pH} = 1$
10. What is the pH of a 0.001 M solution of NaOH in water? $\text{pH} = 11$

1. $\text{H}^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+$ 2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 3. $\text{HCl} \rightleftharpoons \text{H}^+ + \text{Cl}^-$ 4. $\text{NaOH} \rightleftharpoons \text{Na}^+ + \text{OH}^-$ 5. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 6. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 7. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 8. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 9. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 10. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 11. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 12. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 13. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 14. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 15. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 16. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 17. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 18. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 19. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ 20. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$

19 2 Hydrogen Ions And Acidity Section Review Answers

Linda Swisher, Gary A. Thibodeau



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