

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and calculate  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and calculate  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                                    |
|-----------------|--------------------------------------|------------------------------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization constant ( $K_a$ ) |
| acidic solution | acidic solution                      | acidic solution                    |
| acidic solution | acidic solution                      | acidic solution                    |

## Key Equations

- |                                 |                                  |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$               | $pOH = -\log[OH^-]$              |

## Part A Completion

Write the appropriate words in the blanks to complete the sentences. Use the words in the box to complete the sentences.

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
2. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
3. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
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8. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
9. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
10. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .

1.  $1.0 \times 10^{-4}$
2.  $1.0 \times 10^{-10}$
3.  $1.0 \times 10^{-4}$
4.  $1.0 \times 10^{-10}$
5.  $1.0 \times 10^{-4}$
6.  $1.0 \times 10^{-10}$
7.  $1.0 \times 10^{-4}$
8.  $1.0 \times 10^{-10}$
9.  $1.0 \times 10^{-4}$
10.  $1.0 \times 10^{-10}$

## 19 2 Section Review Hydrogen Ions And Acidity

**Ying-Ying Zheng**



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