

Chemistry: Atoms First 2e
17: Kinetics
17.5: Collision Theory

50. Chemical reactions occur when reactants collide. What are two factors that may prevent a collision from producing a chemical reaction?

Solution

The reactants either may be moving too slowly to have enough kinetic energy to exceed the activation energy for the reaction, or the orientation of the molecules when they collide may prevent the reaction from occurring.

52. What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

Solution

The activation energy is the minimum amount of energy necessary to form the activated complex in a reaction. It is usually expressed as the energy necessary to form one mole of activated complex.

54. Describe how graphical methods can be used to determine the activation energy of a reaction from a series of data that includes the rate of reaction at varying temperatures.

Solution

After finding k at several different temperatures, a plot of $\ln k$ versus $\frac{1}{T}$ gives a straight line with the slope $-\frac{E_a}{R}$, from which E_a may be determined.

56. The rate of a certain reaction doubles for every 10 °C rise in temperature.

(a) How much faster does the reaction proceed at 45 °C than at 25 °C?

(b) How much faster does the reaction proceed at 95 °C than at 25 °C?

Solution

(a) The rate doubles for each 10 °C rise in temperature, 45 °C is a 20 °C increase over 25 °C. Thus, the rate doubles two times, or 2^2 (rate at 25 °C) = 4-times faster. (b) 95 °C is a 70 °C increase over 25 °C. Thus the rate doubles seven times, or 2^7 (rate at 25 °C) = 128-times faster.

58. The rate constant at 325 °C for the decomposition reaction $\text{C}_2\text{H}_6 \rightarrow 2\text{C}_2\text{H}_4$ is $6.1 \times 10^{-5} \text{ s}^{-1}$, and the activation energy is 261 kJ per mole of C_2H_6 . Determine the frequency factor for the reaction.

Solution

The rate constant k is related to the activation energy E_a by a relationship known as the Arrhenius equation. Its form is:

$$k = A \times 10^{-(E_a / 2.48 \times 10^4)} = A \times e^{-(E_a / RT)}$$

where A is the frequency factor. Using the data provided, and converting kilojoules to joules:

$$6.1 \times 10^{-5} \text{ s}^{-1} = A \times 10^{-(261,000 \text{ J} / 2.48 \times 10^4 \text{ J K}^{-1} \times 325 \text{ K})}$$

$$= A \times 10^{-22.8}$$

$$A = \frac{6.1 \times 10^{-5} \text{ s}^{-1}}{1.58 \times 10^{-23}} = 3.9 \times 10^{18} \text{ s}^{-1}$$

60. An elevated level of the enzyme alkaline phosphatase (ALP) in human serum is an indication of possible liver or bone disorder. The level of serum ALP is so low that it is very difficult to measure directly. However, ALP catalyzes a number of reactions, and its relative concentration

Sol Mn Chem Exper Theory 2e

Emilie Sanchez



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