

S E R V I C E M A N U A L

SPECTRUM ANALYZER

3585A

VOLUME I



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web a write a balanced equation for the reaction $\text{al}_2\text{o}_3 + 4\text{al} + 3\text{o}_2 \rightarrow 2\text{al}_2\text{o}_3$ b determine the theoretical yield of al_2o_3 160.0 g $\text{al} \times 1 \text{ mol} / 4 \text{ mol} = 0.25 \text{ mol}$ $\text{al}_2\text{o}_3 \times 101.964 \text{ g/mol} = 25.485 \text{ g}$ $\text{al}_2\text{o}_3 \times 302.3 \text{ g/mol} = 75.578 \text{ g}$ $\text{al}_2\text{o}_3 \times 26.982 \text{ g/mol} = 6.745 \text{ g}$ $\text{al} \times 4 \text{ mol} / 1 \text{ mol} = 4 \text{ mol}$ $\text{al}_2\text{o}_3 \times 160.0 \text{ g/mol} = 640.0 \text{ g}$ c determine the percent yield

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

7 2 limiting reagent and reaction yields chem 1114 - Jul 02 2022

web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yieldactual yield 100 based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

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iodine v oxide i_2O_5 reacts with 28.0 grams of carbon monoxide CO to determine the mass of iodine I_2 which could be produced
80 g i_2O_5 1 mol i_2O_5 1 mol I_2 \times 1 333.8 g i_2O_5 1 mol i_2O_5 28 g CO 1 mol CO

7.3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7.3