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# Safeway Turkey Recipe

**Michael Hoelscher**

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2  $\text{CO}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g}$   $\Delta h = 5 \text{ kJ}$   
2  $\text{C}_2\text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g}$   $\Delta h = 2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2 \text{ g} \rightarrow \text{H}_2 \text{ g}$   $\Delta h = 1/2 \text{ kJ}$   
 $\text{O}_2 \text{ g} \rightarrow \text{O}_2 \text{ g}$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + \text{C}_2\text{H}_2 \text{ g} + \text{H}_2 \text{ g}$   $\Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2 \text{ g}$   $\Delta h = 1560 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ g} + \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + \text{H}_2 \text{ g}$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6 \text{ g} + \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + \text{H}_2 \text{ g}$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2 \text{ g} + \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O} \text{ l}$   $\Delta h = 285 \text{ kJ}$  mole

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and  $\Delta h$

values to determine the heat of reaction for the reaction c 2 h 4

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web thus for reaction 9 2 in which 1 mole of h3o and 1 mole of oh combine to form 2 moles of h2o h 55 8 kj as shown note that the enthalpy change h hproducts hreactants is positive if heat is absorbed that is if hproducts

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