

## 16.5

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                              |                    |
|-------------------|------------------------------|--------------------|
| 1. autoionization | 2. conjugate acid of a base  | 3. pH              |
| 4. autoionization | 5. conjugate base of an acid | 6. polyprotic acid |
|                   | 7. weak acid                 |                    |

## Key Equations

- |                                    |  |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Complete the reaction.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acidic hydrogen transfer)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acidic hydrogen transfer)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acidic hydrogen transfer)
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9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
10.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

# 19 2 Hydrogen Ions And Acidity Section Review Answers

**David Baud**



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