

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_m$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by v_f^* (so for most molecules at normal temperatures $v_f^* = 2$ for linear molecules, and 3 for non-linear molecules), then

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Rachel S Tattersall

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Forthcoming Books Rose Arny, 2003 *An Introduction to Physical Science* James T. Shipman, Jerry D. Wilson, Aaron W. Todd, 2002-07-08 This text presents a survey of the physical sciences physics chemistry astronomy meteorology and geology for non science majors. Topics are treated both descriptively and quantitatively providing flexibility for instructors who wish to emphasize a highly descriptive approach a highly quantitative approach or any spectrum in between.

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 - The Rise of Digital Reading Physical Chemistry 9th Edition Instructors Solution Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry 9th Edition Instructors Solution Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - User-Friendly Interface
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- Highlighting and Note-Taking Physical Chemistry 9th Edition Instructors Solution Guide
- Interactive Elements Physical Chemistry 9th Edition Instructors Solution Guide

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Chemistry 9th Edition Instructors Solution Guide

9. Balancing eBooks and Physical Books Physical Chemistry 9th Edition Instructors Solution Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Physical Chemistry 9th Edition Instructors Solution Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Physical Chemistry 9th Edition Instructors Solution Guide

- Setting Reading Goals Physical Chemistry 9th Edition Instructors Solution Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Physical Chemistry 9th Edition Instructors Solution Guide

- Fact-Checking eBook Content of Physical Chemistry 9th Edition Instructors Solution Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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