

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b .
3. Calculate the percent ionization of weak acids.

Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant (K_a)
- acid ionization

- base ionization
- base ionization constant (K_b)

Key Equations

- 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)

- 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2)

Part A Completion

Write the appropriate formula or name for each compound, and write its acid-base behavior in water. Write the K_a or K_b expression for each acid or base.

1. CH_3COOH (acetic acid) $K_a = 1.8 \times 10^{-5}$

2. HNO_3 (nitric acid) $K_a = 2.4 \times 10^{-1}$

3. H_2SO_4 (sulfuric acid) $K_a = 1.0 \times 10^{-1}$

4. H_2CO_3 (carbonic acid) $K_a = 4.3 \times 10^{-7}$

5. $H_2PO_4^-$ (dihydrogen phosphate) $K_a = 6.3 \times 10^{-8}$

6. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

7. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

8. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

9. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

10. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

11. $H_2PO_4^-$ (dihydrogen phosphate) $K_b = 1.6 \times 10^{-7}$

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Marco Cascella



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