

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. acid rain |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the pH of pure water at 25°C? What is the pOH of pure water at 25°C? What is the pH of pure water at 25°C?

What is the pH of a 0.1 M solution of hydrochloric acid? What is the pH of a 0.1 M solution of sodium hydroxide? What is the pH of a 0.1 M solution of acetic acid?

What is the pH of a 0.1 M solution of acetic acid? What is the pH of a 0.1 M solution of acetic acid? What is the pH of a 0.1 M solution of acetic acid? What is the pH of a 0.1 M solution of acetic acid?

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**Kevin T. Patton, Gary A.
Thibodeau, Andrew Hutton**



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