

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| | 7. acid rain | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is appropriate or required.

1. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____ in _____ solution.
2. _____, which concentration is _____, is a _____ acid.
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19 2 Hydrogen Ions And Acidity Section Review Answers

American Society for Testing Materials

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the American Society for Testing Materials text.

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Thibodeau, 2013-01-01 Rev ed of The human body in health disease Gary A Thibodeau Kevin T Patton 5th ed c2010

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