

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**
- k. **H₂SO₄**: **hydrobromic acid**
- l. **H₂SO₄**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

| Acid | Base | Mixture | Neutral | Acidic | Basic | Neutral | Basic |
|------|------|---------|---------|--------|-------|---------|-------|
|------|------|---------|---------|--------|-------|---------|-------|

3. Write the dissociation for **H₂SO₄**.

4. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because HClO_4 is a strong acid.

c. Dissolve **CaCl₂** in a portion of water and a portion of vinegar? **common** because it is strong at $\text{pH} \approx 7$.

d. In your pH -response, is **HClO₄** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate base of **HClO₄**: **ClO₄⁻**

f. Identify **HClO₄** as aprotic, amphotropic, or polyprotic. **polyprotic** because it can lose multiple H^+ ions.

5. **H₂SO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{SO}_4^{2-}$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because **H₂SO₄** is a strong acid.

c. Dissolve **CaCl₂** in a portion of water and a portion of vinegar? **common** because it is strong at $\text{pH} \approx 7$.

d. In your pH -response, is **H₂SO₄** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate acid of **H₂SO₄**: **HSO₄⁻**

f. Identify **H₂SO₄** as aprotic, amphotropic, or polyprotic. **polyprotic**

6. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because **HClO₄** is a strong acid.

c. Dissolve **CaCl₂** in a portion of water and a portion of vinegar? **common** because it is strong at $\text{pH} \approx 7$.

d. In your pH -response, is **HClO₄** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid** because it is a strong acid.

e. Identify the conjugate base of **HClO₄**: **ClO₄⁻**

f. Identify **HClO₄** as aprotic, amphotropic, or polyprotic. **polyprotic** because it can lose multiple H^+ ions in the stronger acid.

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J Dewey

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