

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|-----------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or hydroxide ion concentration of the solution and then the pH or pOH of the solution. Round answers to the correct number of significant figures.

1. What substance is _____, contains hydrogen ions?
2. _____ is the concentration of those acids present in rainwater that equal to _____.
3. _____ is the concentration of those acids present in _____.
4. _____ is the concentration of those acids present in _____.
5. _____ is the concentration of those acids present in _____.
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Velavancorresponding



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