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Fuse-Box info



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# 2004 Bmw X5 Fuse Box

**Michael Seilmaier**

A red circular graphic with a gradient, appearing as a stylized arrow or a light effect, positioned to the right of the author's name.

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$  text cu 2 text agno<sub>3</sub> rightarrow 2 text ag text cu no<sub>3</sub> 2

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0 g Al × 1 mol Al / 27.0 g Al × 2 mol Al<sub>2</sub>O<sub>3</sub> / 3 mol Al = 119.6 g Al<sub>2</sub>O<sub>3</sub> c determine the percent yield 160.0 g / 119.6 g × 100 = 134.2%

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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