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# 2004 Bmw X5 Fuse Box

**Michael Seilmaier**

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 $4 Cu + 2 AgNO_3 \rightarrow 2 Cu(NO_3)_2 + 2 Ag$

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0 g Al × 1 mol Al × 2 mol Al<sub>2</sub>O<sub>3</sub> / 101.964 g Al<sub>2</sub>O<sub>3</sub> = 302.3 g Al<sub>2</sub>O<sub>3</sub> c determine the percent yield = actual yield / theoretical yield × 100 = 260.0 g / 302.3 g = 86.01%

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web 7 Na + 2C + 2O → 4 is the limiting reactant percent yield = 86.6% 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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