

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element in the formula.

1. Acetic acid, CH_3COOH
2. Hydrofluoric acid, HF
3. Hydrochloric acid, HCl
4. Hydrobromic acid, HBr
5. Hydroiodic acid, HI
6. Nitric acid, HNO_3
7. Sulfuric acid, H_2SO_4
8. Phosphoric acid, H_3PO_4
9. Acetic acid, CH_3COOH
10. Hydrofluoric acid, HF
11. Hydrochloric acid, HCl
12. Hydrobromic acid, HBr
13. Hydroiodic acid, HI
14. Nitric acid, HNO_3
15. Sulfuric acid, H_2SO_4
16. Phosphoric acid, H_3PO_4

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Valerie C. Scanlon, Tina Sanders



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