

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                            |                   |              |
|----------------------------|-------------------|--------------|
| acid dissociation constant | 1. conjugate base | 2. weak acid |
| acid ionization            | 3. conjugate acid | 3. weak base |
| acid ionization            | 4. conjugate base | 4. weak acid |

## Key Equations

- |                                    |                                     |
|------------------------------------|-------------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| 3. $pH = -\log[H^+]$               | 4. $pOH = -\log[OH^-]$              |

## Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Write the chemical formula of the acid or base in the blank.

1. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ . The  $K_b$  of acetate ion is  $5.6 \times 10^{-10}$ .
2. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ . The  $K_a$  of ammonium ion is  $5.6 \times 10^{-10}$ .
3. The  $K_a$  of hydrofluoric acid is  $6.6 \times 10^{-4}$ . The  $K_b$  of fluoride ion is  $1.5 \times 10^{-11}$ .
4. The  $K_b$  of pyridine is  $1.7 \times 10^{-9}$ . The  $K_a$  of pyridinium ion is  $5.9 \times 10^{-10}$ .
5. The  $K_a$  of formic acid is  $1.8 \times 10^{-4}$ . The  $K_b$  of formate ion is  $5.6 \times 10^{-11}$ .
6. The  $K_b$  of methylamine is  $4.4 \times 10^{-4}$ . The  $K_a$  of methylammonium ion is  $2.3 \times 10^{-11}$ .

7. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
8. The  $K_b$  of acetate ion is  $5.6 \times 10^{-10}$ .
9. The  $K_a$  of hydrofluoric acid is  $6.6 \times 10^{-4}$ .
10. The  $K_b$  of fluoride ion is  $1.5 \times 10^{-11}$ .
11. The  $K_a$  of formic acid is  $1.8 \times 10^{-4}$ .
12. The  $K_b$  of formate ion is  $5.6 \times 10^{-11}$ .
13. The  $K_a$  of pyridinium ion is  $5.9 \times 10^{-10}$ .
14. The  $K_b$  of pyridine is  $1.7 \times 10^{-9}$ .
15. The  $K_a$  of methylammonium ion is  $2.3 \times 10^{-11}$ .
16. The  $K_b$  of methylamine is  $4.4 \times 10^{-4}$ .

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**D Keegan**



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