

Robert Smith

Staff Payroll Accountant

PERSONAL STATEMENT

Detailed oriented Payroll Accountant with over 4 years of experience in the Engineering domain within large, high-volume organizations, utilizing leading technologies and processes. Analytical and highly-adaptable, skilled in working across all business and operational groups and leadership levels.

WORK EXPERIENCE

Staff Payroll Accountant **Sparrows Offshore - March 2013 - Present**

Responsibilities:

- Collecting daily, weekly and bi-weekly timesheets from office, field, onshore and offshore employees.
- Overseeing the time and labor processes to ensure timesheets are reconciled to field tickets by creating workflows with offshore account managers to verify accuracy and completeness.
- Primary point of contact for all payroll, tax, and general ledger reconciling items, ensuring replies.
- Producing weekly reports for Payroll, Finance, and HR, and our international payroll partners.
- Preparing payroll journal entries, account reconciliations, expense journals, and accruals accounts.
- Handling the general ledger, interface mapping, and balance sheet accounts, and file related.
- Keeping U.S HRIS system updated, maintaining and file timesheets, payroll, and accounting data both paper and electronically.

PAYROLL SPECIALIST **Exterran Energy Solutions - September 2014 - March 2015**

Responsibilities:

- Audited all payroll setups, i.e., tables, background calculations, pay and deduction codes.
- Sustained quarterly payroll tax balancing and reconciliations to year-end W2s, W2Cs.
- Filed for new tax account setups with entities, additions, and terminations of payroll tax accounts.
- Managed the addition, termination, and setup of new payroll tax accounts.
- Accurate and timely biweekly payroll processing using the ADP Workforce Now platform for both exempt and non-exempt employees.
- Ensured the compliance with company policies, as well as federal / state payroll laws and regulations.
- This is Dummy Description data, Replace with job description relevant to your current role. In case if you dont need it you can delete it.

CONTACT DETAILS

1737 Marshville Road,
Alabama
(123)-456-7890
info@qwikresume.com
www.qwikresume.com

SKILLS

General Ledger, ADP, HR Information System, Benefits Administration, CITRIX, Reconciliation, Pivot Tables, V-looks Ups

LANGUAGES

English (Native)
French (Professional)
Spanish (Professional)

INTERESTS

Climbing
Snowboarding
Cooking
Reading

REFERENCES

Reference - 1 (Company Name)
Reference - 2 (Company Name)

Accounting Practice Set Answers Payroll

Bruce Reinhart

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the author's name.

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DivisionRecords & Briefs New York State Appellate Division , *The Publishers' Trade List Annual* ,1978 *Human Resources Report* ,2002 *National Business Education Quarterly* ,1947 **Financial Accounting, Self-Study Problems/Solutions Book** Jerry J. Weygandt,Donald E. Kieso,Paul D. Kimmel,1999-11-24 From one of the hottest author teams in accounting comes a new edition of the highly successful Financial Accounting This edition retains the traditional procedural coverage of the previous editions along with a practical decision making focus Anyone who will ever be asked to prepare or use accounting information to make effective decisions will benefit from this best selling text *El-Hi textbooks in print* R. R. Bowker LLC,1983

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2 \text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq} + 2 \text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2 \text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3 \text{Ba}^{2+} \text{aq} + 2 \text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ net ionic $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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web the $\text{ag}^+ \text{aq}$ and $\text{cl}^- \text{aq}$ ions become agcl s but the $\text{na}^+ \text{aq}$ ions and the $\text{no}_3^- \text{aq}$ ions stay as $\text{na}^+ \text{aq}$ ions and $\text{no}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{cl}^- \text{aq} + \text{cu}^{2+} \text{aq} \rightarrow \text{cucl}_2 \text{s}$ $2\text{cl}^- \text{aq} + \text{cu}^{2+} \text{aq} \rightarrow \text{cucl}_2 \text{s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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