

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

## Vocabulary

- |                   |   |                  |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion ( $\text{H}_3\text{O}^+$ ) | 3. acid solution |
| 4. autoionization | 5. acid rain                                | 6. acid rain     |
| 7. autoionization | 8. acid rain                                | 9. pH            |

## Key Equation

- |                                    |  |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

## Part A Completion

Write the equation for the autoionization of water. What is the concentration of  $\text{H}^+$  in pure water? What is the concentration of  $\text{OH}^-$  in pure water?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

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What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{H}^+$ ?

What is the pH of a solution with a concentration of  $1.0 \times 10^{-7} \text{ M}$   $\text{OH}^-$ ?

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

2.  $1.0 \times 10^{-7} \text{ M}$

3.  $1.0 \times 10^{-7} \text{ M}$

4.  $1.0 \times 10^{-7} \text{ M}$

5.  $1.0 \times 10^{-7} \text{ M}$

6.  $1.0 \times 10^{-7} \text{ M}$

7.  $1.0 \times 10^{-7} \text{ M}$

8.  $1.0 \times 10^{-7} \text{ M}$

9.  $1.0 \times 10^{-7} \text{ M}$

10.  $1.0 \times 10^{-7} \text{ M}$

11.  $1.0 \times 10^{-7} \text{ M}$

12.  $1.0 \times 10^{-7} \text{ M}$

13.  $1.0 \times 10^{-7} \text{ M}$

14.  $1.0 \times 10^{-7} \text{ M}$

15.  $1.0 \times 10^{-7} \text{ M}$

16.  $1.0 \times 10^{-7} \text{ M}$

17.  $1.0 \times 10^{-7} \text{ M}$

18.  $1.0 \times 10^{-7} \text{ M}$

19.  $1.0 \times 10^{-7} \text{ M}$

20.  $1.0 \times 10^{-7} \text{ M}$

# 19 2 Hydrogen Ions And Acidity Answer Key

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