

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pOH | 6. acid solution |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below. Complete each equation as needed.

1. When $[\text{H}^+] = 1.0 \times 10^{-4} \text{ M}$, the autoionization constant is 1.0×10^{-14} .
2. When $[\text{H}^+] = 1.0 \times 10^{-4} \text{ M}$, the autoionization constant is 1.0×10^{-14} .
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20. When $[\text{H}^+] = 1.0 \times 10^{-4} \text{ M}$, the autoionization constant is 1.0×10^{-14} .

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19 2 Hydrogen Ions And Acidity Worksheet Answers

Rod O'Connor



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