



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ion concentration

### Points of Consideration

What are the components of acids and bases? How do acids and bases affect the pH of a solution? How do acids and bases affect the solubility of other substances?

What is the difference between a strong acid and a weak acid? How do the concentrations of strong acids and weak acids differ?

What is the difference between a strong base and a weak base? How do the concentrations of strong bases and weak bases differ?

What is the difference between an acid and a base? How do the properties of acids and bases differ?

What is the difference between a strong acid and a weak acid?

### Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = 7$	4. $\text{pH} < 7$
5. $\text{pH} > 7$	6. $\text{pH} = 7$
7. $\text{pH} = 7$	8. $\text{pH} < 7$
9. $\text{pH} > 7$	10. $\text{pH} = 7$
11. $\text{pH} < 7$	12. $\text{pH} > 7$
13. $\text{pH} = 7$	14. $\text{pH} < 7$
15. $\text{pH} > 7$	16. $\text{pH} = 7$
17. $\text{pH} < 7$	18. $\text{pH} > 7$
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*Effluent & Water Treatment Journal* ,1983 **Physics Briefs** ,1989-07

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## 19 2 Hydrogen Ions And Acidity Answer Key Introduction

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