

Chemistry Study Guide

| Periodic Table of the Elements | | | | | | | | | | | | | | | | | | |
|--------------------------------|----|----|----|----|----|----|----|----|----|----|-----|----|-----|----|-----|-----|----|--|
| H | Li | Be | | | | | | | | | | | | | | | | |
| Na | Mg | | | | | | | | | | | | | | | | | |
| K | Ca | Sc | Ti | V | Cr | Mn | Fe | Co | Ni | Cu | Zn | Ga | Ge | As | Se | Br | Kr | |
| Rb | Sr | Y | Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag | Cd | In | Sn | Sb | Te | I | Xe | |
| Cs | Ba | Hf | Ta | W | Re | Os | Ir | Pt | Au | Hg | Tl | Pb | Bi | Po | At | Rn | | |
| Fr | Ra | Rf | Db | Sg | Bh | Hs | Mt | Ds | Rg | Cn | Uut | Fl | Uup | Lv | Uus | Uuo | | |
| | | La | Ce | Pr | Nd | Pm | Sm | Eu | Gd | Tb | Dy | Ho | Er | Tm | Yb | Lu | | |
| | | Ac | Th | Pa | U | Np | Pu | Am | Cm | Bk | Cf | Es | Fm | Md | No | Lr | | |

Objectives:

To understand the structure of an atom, and the nature + properties of subatomic particles.



Matter:

Anything that has mass and occupies space.

Examples: air, gases, liquids, solidz!

Atoms:

Atoms are the smallest, indivisible particle.

- Outer part → shell/orbit/energy level → carries the electronz (e-)
- Inner part → nucleus → houses the protonz (p+) and neutronz (n)
- Outer & Inner part → together these are the subatomic particles.

Subatomic Particle Table:

| Particle | Symbol | Mass (amu) | Charge | Location |
|----------|----------------|------------|--------|-------------|
| Proton | p ⁺ | 1 | +1 | Nucleus |
| Neutron | n | 1 | 0 | Nucleus |
| Electron | e ⁻ | -0 | -1 | Orbit/Shell |

Section 2 nuclear Chemistry Study Guide Key

Darrell D. Ebbing

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