

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate formula or name for each of the following acids and bases, showing the chemical formula and name of the acid or base.

1. A weak acid with  $K_a = 1.0 \times 10^{-4}$  is  $CH_3COOH$ .
2. A weak base with  $K_b = 1.0 \times 10^{-4}$  is  $CH_3NH_2$ .
3. A weak acid with  $K_a = 1.0 \times 10^{-4}$  is  $CH_3COOH$ .
4. A weak base with  $K_b = 1.0 \times 10^{-4}$  is  $CH_3NH_2$ .
5. A weak acid with  $K_a = 1.0 \times 10^{-4}$  is  $CH_3COOH$ .
6. A weak base with  $K_b = 1.0 \times 10^{-4}$  is  $CH_3NH_2$ .
7. A weak acid with  $K_a = 1.0 \times 10^{-4}$  is  $CH_3COOH$ .
8. A weak base with  $K_b = 1.0 \times 10^{-4}$  is  $CH_3NH_2$ .
9. A weak acid with  $K_a = 1.0 \times 10^{-4}$  is  $CH_3COOH$ .
10. A weak base with  $K_b = 1.0 \times 10^{-4}$  is  $CH_3NH_2$ .

1.  $CH_3COOH$
2.  $CH_3NH_2$
3.  $CH_3COOH$
4.  $CH_3NH_2$
5.  $CH_3COOH$
6.  $CH_3NH_2$
7.  $CH_3COOH$
8.  $CH_3NH_2$
9.  $CH_3COOH$
10.  $CH_3NH_2$

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**Yicheng Fang**



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