



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base in terms of the Arrhenius theory.
- Define strong and weak acids.

Definitions

- Hydrogen ion
- Hydrogen ion concentration
- Acid
- Base
- Strong acid
- Weak acid
- Strong base
- Weak base

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Principles of Acids and Bases

Acids and bases are defined in terms of their ability to ionize in aqueous solution. Acids are proton donors, and bases are proton acceptors.

Strong acids dissociate completely in aqueous solution. The concentration of these acids is equal to the concentration of their ions.

Weak acids dissociate partially in aqueous solution. The concentration of the acid is greater than the concentration of its ions.

Strong bases dissociate completely in aqueous solution. The concentration of the base is equal to the concentration of its ions.

Weak bases dissociate partially in aqueous solution.

Strong Acids

- Hydrochloric acid
- Sulfuric acid
- Acetic acid
- Hydrobromic acid
- Hydriodic acid
- Hydrofluoric acid
- Hydrocyanic acid
- Hydroxylamine
- Hydrazoic acid
- Hydrazine
- Hydrazine sulfate
- Hydrazine hydrochloride
- Hydrazine hydrobromide
- Hydrazine hydroiodide
- Hydrazine hydrofluoride
- Hydrazine hydrocyanide

19 2 Hydrogen Ions And Acidity Answer Key

John Arthur Wilson

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