

## Solving Quadratic Equations by Factoring Practice

Date \_\_\_\_\_

Solve each equation by factoring.

1)  $x^2 + 6x + 9 = 0$

- A)  $\{-3, -5\}$       B)  $\{3\}$   
C)  $\{-6, 3\}$       D)  $\{-3\}$

2)  $n^2 - 7n = 0$

- A)  $\{-7, -3\}$       B)  $\{12\}$   
C)  $\{7, 0\}$       D)  $\{-12, 7\}$

3)  $n^2 - 2n - 3 = 0$

- A)  $\{3, -1\}$       B)  $\{-3, 1\}$   
C)  $\{-1, 5\}$       D)  $\{-8, -10\}$

4)  $a^2 - 11a + 30 = 0$

- A)  $\{5, -6\}$       B)  $\{8, -9\}$   
C)  $\{-6, -9\}$       D)  $\{5, 6\}$

5)  $b^2 + 6b - 40 = 0$

- A)  $\{1, -10\}$       B)  $\{3, -10\}$   
C)  $\{-8, -10\}$       D)  $\{4, -10\}$

6)  $x^2 - 5x - 66 = 0$

- A)  $\{-2, -6\}$       B)  $\{11, -6\}$   
C)  $\{-8, 9\}$       D)  $\{11, -2\}$

7)  $n^2 - 19n + 88 = 0$

- A)  $\{-4, 11\}$       B)  $\{-4, 8\}$   
C)  $\{8, 11\}$       D)  $\{5, -11\}$

8)  $x^2 + 6x - 27 = 0$

- A)  $\{4, -7\}$       B)  $\{-9, -11\}$   
C)  $\{-11, 9\}$       D)  $\{3, -9\}$

# Tranforming Quadratics Kuta

**Hussin A.Rothana**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{NaOH}$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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