

**Chemistry 2<sup>nd</sup> Semester Exam Review Sheet:****Chapters 9, 10, 11, 12, 13, 14, 15, and 16****Chapter 9: Chemical Quantities**

1. Solve stoichiometry problems using grams  $\rightarrow$  moles  $\rightarrow$  moles  $\rightarrow$  grams calculations. Examples below:  
a. Calculate moles/grams of  $(\text{NH}_4)_2\text{SO}_4$  produced when 30.0 moles of  $\text{NH}_3$  react with  $\text{H}_2\text{SO}_4$ .



$$30.0 \text{ moles} \times 1 \text{ mol/2 mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ g/mol/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b.  $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$ . How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams  $\rightarrow$  Moles  $\rightarrow$  Moles  $\rightarrow$  Grams

$$100. \text{ g } \text{Na}_2\text{CO}_3 \times 1 \text{ mol/106.0 g} \times 2 \text{ mol } \text{NaNO}_3 / 1 \text{ mol } \text{Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g } \text{NaNO}_3 \times 1 \text{ mol/85.00 g} \times 2 \text{ mol / 2 mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of  $\text{NaNO}_3$  if 100.0 grams of  $\text{Na}_2\text{CO}_3$  are used in the reaction? What is the percent yield of  $\text{NaNO}_3$  if 130.5 grams of  $\text{NaNO}_3$  are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

**Chapter 10: Modern Atomic Theory**

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? Absorb a photon of energy.

How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms (page 285-86)

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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**Source Book for Chemistry Teachers** William Thomas Lippincott,1981      *Engineering Education* ,1983      **Cornell University Courses of Study** Cornell University,2003      *THE Journal* ,1984      Catalog Issue University of Colorado,1974

**Summer School Number** Southwestern at Memphis,1933      **University of Colorado at Denver Catalog** University of Colorado at Denver,2009      **Current Index to Journals in Education** ,1986      Proceedings ,1984      *The School Review* ,1914      **Computer Aided Training in Science and Technology** E. Oñate,1990      **Science Books & Films** ,1980      **American Journal of Physics** ,1994      **Current Index to Journals in Education, Semi-Annual Cumulation, July-December, 1976** ,1976-10      **Abridged Index Medicus** ,1976      Self-study Report: Self-study report : appendices -v. 3. Abstract of the Self-study report University of Illinois at Urbana-Champaign,1979      **Computer-assisted Instruction** Albert E. Hickey,1968      University of Kentucky Catalogue University of Kentucky,1973      Paperbound Books in Print ,1966      **Resources in Education** ,1978

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