

# Chemistry Study Guide

# Periodic Table of the Elements

<small>1</small> <b>H</b> <small>1.008</small>																	<small>2</small> <b>He</b> <small>4.003</small>																															
<small>3</small> <b>Li</b> <small>6.941</small>	<small>4</small> <b>Be</b> <small>9.012</small>																	<small>5</small> <b>B</b> <small>10.81</small>	<small>6</small> <b>C</b> <small>12.01</small>	<small>7</small> <b>N</b> <small>14.01</small>	<small>8</small> <b>O</b> <small>15.99</small>	<small>9</small> <b>F</b> <small>18.99</small>	<small>10</small> <b>Ne</b> <small>20.18</small>																									
<small>11</small> <b>Na</b> <small>22.99</small>	<small>12</small> <b>Mg</b> <small>24.31</small>																	<small>13</small> <b>Al</b> <small>26.98</small>	<small>14</small> <b>Si</b> <small>28.09</small>	<small>15</small> <b>P</b> <small>30.97</small>	<small>16</small> <b>S</b> <small>32.06</small>	<small>17</small> <b>Cl</b> <small>35.45</small>	<small>18</small> <b>Ar</b> <small>39.95</small>																									
<small>19</small> <b>K</b> <small>39.09</small>	<small>20</small> <b>Ca</b> <small>40.08</small>	<small>21</small> <b>Sc</b> <small>44.96</small>	<small>22</small> <b>Ti</b> <small>47.88</small>	<small>23</small> <b>V</b> <small>50.94</small>	<small>24</small> <b>Cr</b> <small>51.99</small>	<small>25</small> <b>Mn</b> <small>54.94</small>	<small>26</small> <b>Fe</b> <small>55.85</small>	<small>27</small> <b>Co</b> <small>58.93</small>	<small>28</small> <b>Ni</b> <small>58.71</small>	<small>29</small> <b>Cu</b> <small>63.55</small>	<small>30</small> <b>Zn</b> <small>65.38</small>	<small>31</small> <b>Ga</b> <small>69.72</small>	<small>32</small> <b>Ge</b> <small>72.64</small>	<small>33</small> <b>As</b> <small>74.92</small>	<small>34</small> <b>Se</b> <small>78.96</small>	<small>35</small> <b>Br</b> <small>79.90</small>	<small>36</small> <b>Kr</b> <small>83.80</small>																															
<small>37</small> <b>Rb</b> <small>85.47</small>	<small>38</small> <b>Sr</b> <small>87.62</small>	<small>39</small> <b>Y</b> <small>88.91</small>	<small>40</small> <b>Zr</b> <small>91.22</small>	<small>41</small> <b>Nb</b> <small>92.91</small>	<small>42</small> <b>Mo</b> <small>95.94</small>	<small>43</small> <b>Tc</b> <small>98.91</small>	<small>44</small> <b>Ru</b> <small>101.07</small>	<small>45</small> <b>Rh</b> <small>101.07</small>	<small>46</small> <b>Pd</b> <small>106.32</small>	<small>47</small> <b>Ag</b> <small>107.87</small>	<small>48</small> <b>Cd</b> <small>112.41</small>	<small>49</small> <b>In</b> <small>114.82</small>	<small>50</small> <b>Sn</b> <small>118.71</small>	<small>51</small> <b>Sb</b> <small>121.76</small>	<small>52</small> <b>Te</b> <small>127.60</small>	<small>53</small> <b>I</b> <small>126.91</small>	<small>54</small> <b>Xe</b> <small>131.29</small>																															
<small>55</small> <b>Cs</b> <small>132.91</small>	<small>56</small> <b>Ba</b> <small>137.33</small>			<small>57</small> <b>Hf</b> <small>178.49</small>	<small>58</small> <b>Ta</b> <small>180.95</small>	<small>59</small> <b>W</b> <small>183.84</small>	<small>60</small> <b>Re</b> <small>186.21</small>	<small>61</small> <b>Os</b> <small>190.23</small>	<small>62</small> <b>Ir</b> <small>192.22</small>	<small>63</small> <b>Pt</b> <small>195.08</small>	<small>64</small> <b>Au</b> <small>196.97</small>	<small>65</small> <b>Hg</b> <small>200.59</small>	<small>66</small> <b>Tl</b> <small>204.38</small>	<small>67</small> <b>Pb</b> <small>207.2</small>	<small>68</small> <b>Bi</b> <small>208.98</small>	<small>69</small> <b>Po</b> <small>209</small>	<small>70</small> <b>At</b> <small>210</small>	<small>71</small> <b>Rn</b> <small>222</small>																														
<small>87</small> <b>Fr</b> <small>223</small>	<small>88</small> <b>Ra</b> <small>226</small>			<small>89</small> <b>Rf</b> <small>261</small>	<small>90</small> <b>Db</b> <small>262</small>	<small>91</small> <b>Sg</b> <small>266</small>	<small>92</small> <b>Bh</b> <small>264</small>	<small>93</small> <b>Hs</b> <small>277</small>	<small>94</small> <b>Mt</b> <small>268</small>	<small>95</small> <b>Ds</b> <small>271</small>	<small>96</small> <b>Rg</b> <small>272</small>	<small>97</small> <b>Cn</b> <small>285</small>	<small>98</small> <b>Uut</b> <small>289</small>	<small>99</small> <b>Fl</b> <small>288</small>	<small>100</small> <b>Uup</b> <small>294</small>	<small>101</small> <b>Lv</b> <small>293</small>	<small>102</small> <b>Uus</b> <small>294</small>	<small>103</small> <b>Uuo</b> <small>294</small>																														
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**Charles E. Mortimer**



## Section 2nuclear Chemistry Study Guide Key:

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